Percent Composition

**Purpose:** During today’s lesson, we will learn about percent composition, a concept that you are likely to be familiar with from math class and from your own life. It is a helpful mathematical concept to use in chemistry, because it allows us to determine just *how much of an element (by weight)* is present in a molecule or a mole of molecules. This knowledge will become increasingly important as we talk about *chemical reactions* and *how much of each reactant* will be needed to produce a certain amount of product. **This worksheet will be turned in the day of the test.**

**Task:** We will be working as a class to build the mathematical formula for percent composition. This will be done through class discussion and individual and pairs-work. Be prepared to use your problem-solving skills to determine a procedure for attacking problems involving percent composition!

## Bellwork:

## General formula for Percent Composition:

## Percent composition of a molecule:

PCl5

|  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- |
|  | Chemical Name | Drawing | Mole fraction  for each element | Molar Mass  for each element | Percent composition  for each element |
| NaCl |  |  |  |  |  |
| FeSO4 |  |  |  |  |  |
| H2SO4 |  |  |  |  |  |