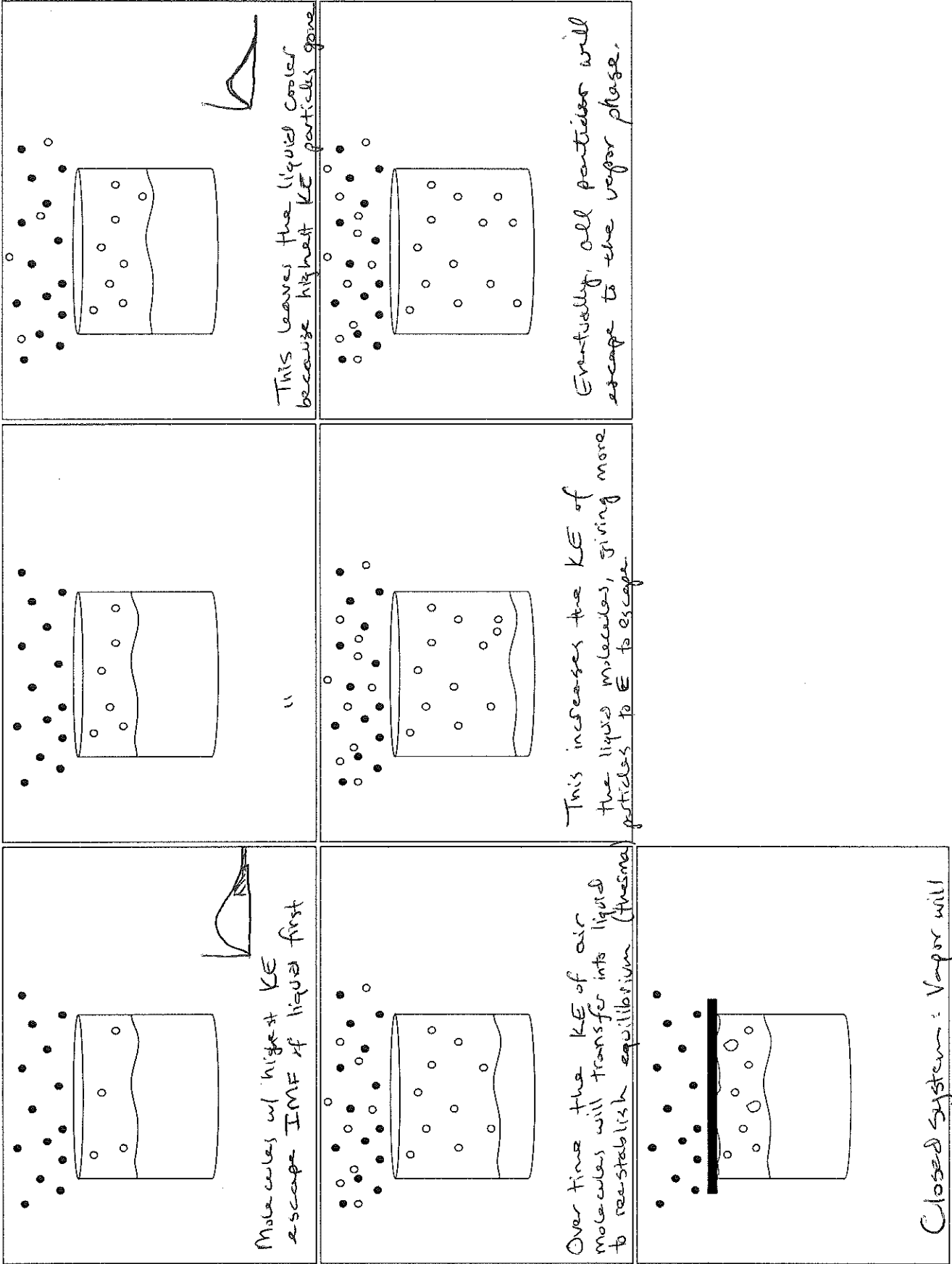
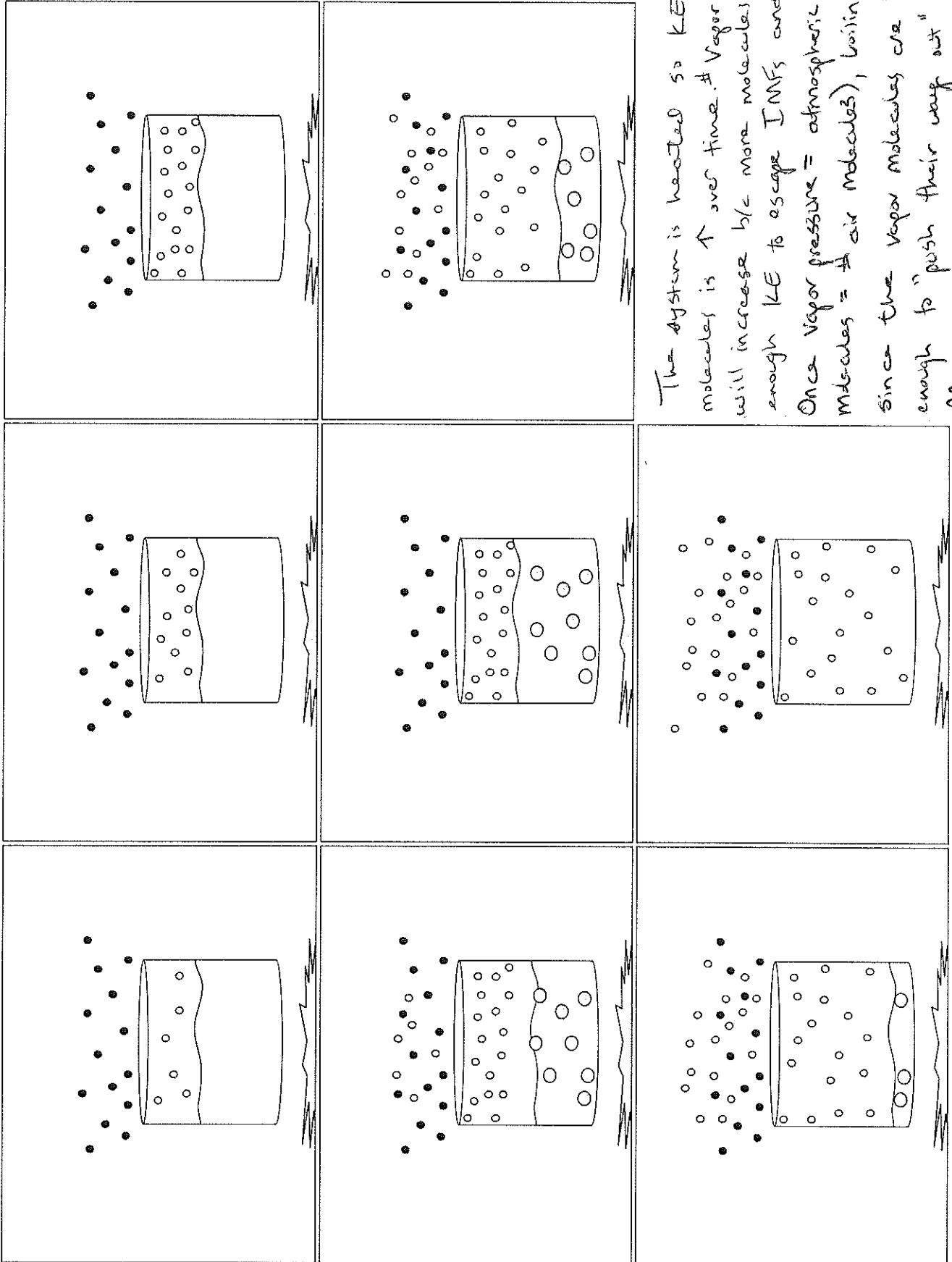


Process of evaporation



Process of boiling a solution



The system is heated so KE of

molecules is \uparrow over time. # Vapor molecules will increase b/c more molecules will have enough KE to escape IMFs and vaporize

Once vapor pressure = atmospheric pressure (# vapor molecules = # air molecules), boiling occurs

since the vapor molecules are now numerous enough to "push their way out" of the beaker.

~~After~~ KE of molecules in liquid will continue to \uparrow due to heat source. Eventually all liquid molecules will vaporize.